ELECTRIC CONDUCTIVITY OF CARBON DIOXIDE AQUEOUS SOLUTIONS

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S u m m a r y

The system CO_2 + H_2O with the concentration of free carbon dioxide varying from those close to the saturation one to the equilibrium one under environmental conditions has been studied. The dependences of ionic component concentrations on the solution pH are determined. Simultaneous measurements of the electric conductivity and the pH of the solution show that the contributions of ionic components in the carbonate-water system to the electric conductivity of the solution are additive and satisfy the Kohlrausch law with an error not exceeding ±0.5%.