

BASIC AND SUBSIDIARY ATOM PARAMETERS.
2. ELECTRONEGATIVITY, BINDING ENERGY,
AND FERSMANS OF ATOMS

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We show that the electronegativity of an atom is connected with its Fersman by the equations of the first and second degrees, and the energy of the cation-anion interaction is connected with the difference of the Fersmans of atoms. Fersman is a complex parameter connected with the serial number of an element, an ion radius, and a summary potential of ionization at a maximum valency.